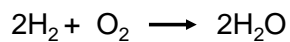


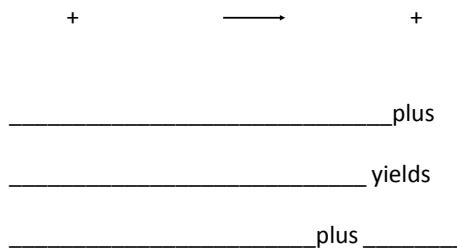
PS Chemistry Notes- Balancing Equations

How matter changes.

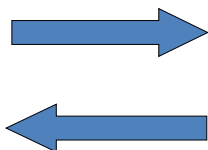
- Chemical Reaction -
- Chemical Equation -



Example:



- Reactant(s) -
- Product(s) -



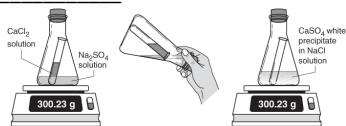
Practice

Reactants Products

$\text{Fe} + \text{S} \longrightarrow \text{FeS}$		
$\text{H}_2\text{SO}_4 + \text{Zn} \longrightarrow \text{ZnSO}_4 + \text{H}_2$		
$\text{Mg} + \text{S} \longrightarrow \text{MgS}$		
$\text{AgNO}_3 + \text{NaCl} \longrightarrow \text{NaNO}_3 + \text{AgCl}$		

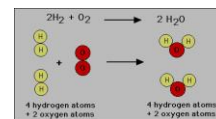
Law of Conservation of Matter

- Matter _____ in a chemical change.
- The same _____ are present before and after the reaction.
- The mass of the _____ equals the mass of the _____

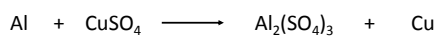
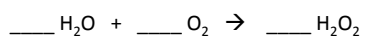
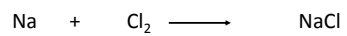


How do you balance the equation?

- Balancing equations satisfies the _____
- The same number and kinds of atom _____
- _____ on both sides of the equation.
- Use _____ to balance.
NOT subscripts! NOT subscripts! NOT subscripts!



Examples



Practice

