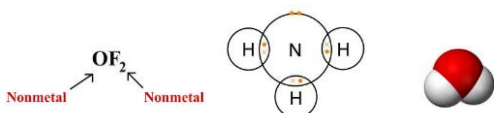


NAMING COVALENT COMPOUNDS

**BINARY COVALENT COMPOUNDS:
COMPOSED OF...**



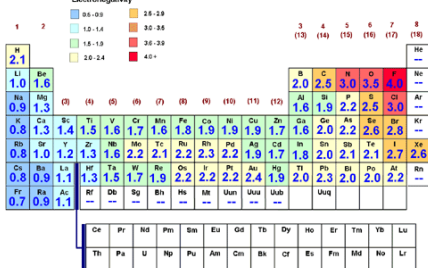
PREFIXES-STAND FOR THE AMOUNT OF ATOMS PRESENT

Number of Atoms	Prefix
1	mono-
2	di-
3	tri-
4	tetra-
5	penta-
6	hexa-
7	hepta-
8	octa-
9	nona-
10	deca-

Steps to Naming Covalent Compounds

1. The element with less electronegativity is written _____ in the name.

Exception: when the compound contains oxygen and a halogen, the name of the halogen is the **first** word in the name.



a. If there is only _____ atom present _____ use a prefix.

b. If there is _____ present then you must use the correct **prefix**.

Examples:

CO₂ Carbon...NOT _____

N₂O₅ _____

2. The second element in the name is named as if it were an anion, i.e., by adding the suffix - _____ to the name of the element.

Add the _____ to indicate how many atoms are present. (including mono).

Note: when the addition of the Greek prefix places two vowels adjacent to one another, the "a" (or the "o") at the end of the Greek prefix is usually dropped to avoid "_____" or "_____" combinations, but not "io".

For example: "nonaoxide" would be written as "_____", and "monoxide" would be written as "_____".

The "i" at the end of the prefixes "_____" and "_____" are never dropped. **For example:** SO₂= _____

(NOT sulfur dioxide)

Examples:

CO₂

N₂O₅

Check your understanding:

Cl₂O

PCl₃

Cl₄