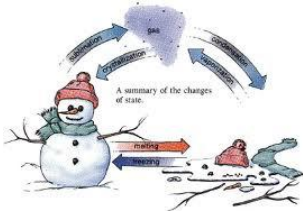


Chapter 16, Section 1



States of Matter

1. Solid
 -
 -
2. Liquid
 -
 -
3. Gas
 -
4. Plasma
 -

Kinetic Energy of Matter

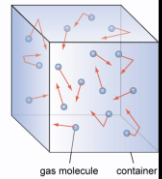
Temperature = _____ of particles (_____ the particles are moving)



- Higher temperature = _____ moving particles
- Lower temperature = _____ moving particles

Kinetic Theory of Matter

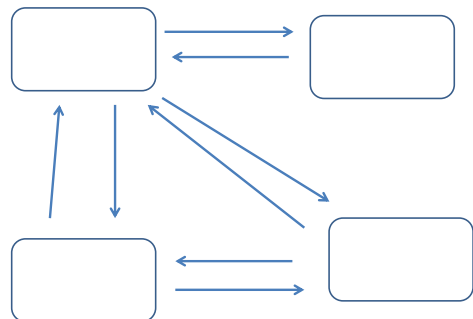
- The particles of matter are in _____.
- _____ causes the particles to move faster (causing _____ with each other and the container)
- The faster they move the more the particles will try to _____.



Behavior of Particles

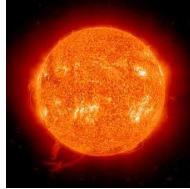
- **Solid**-particles are _____ and are _____.
- **Liquid**-particles are _____, having enough energy to _____
- **Gas**-particles are _____ and _____
- **Plasma**-particle move _____

Changes in States of Matter

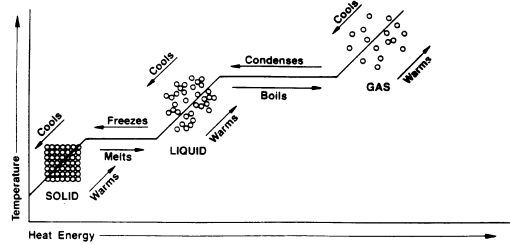


Most common state of matter?

- Scientists estimate that the most common state of matter in the universe is _____.
- Examples:



Phase Change Graph



Unusual materials

Some solids soften gradually over a temperature range (does not have a melting point)



- Examples:

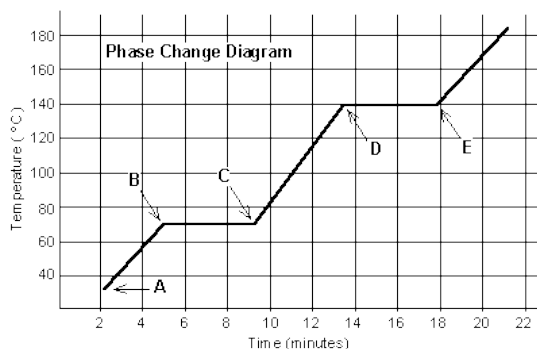
Unusual materials



As it changes from a solid to a liquid, the substance starts to flow BUT it doesn't lose its ordered arrangement

- Examples:

Check Your Understanding



Check Your Understanding

- At **point A** the substance exists as a solid. As time pass minute, heat is added to the substance. What happens to the kinetic energy of the particles?
- At **point B** the substance has absorbed enough energy so that it begins to melt; therefore, at **point C** the substance is in what state of matter?
- Describe what is happening to the substance from 9-13 minutes.
- At **point E** the substance is in what state of matter?