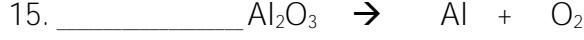
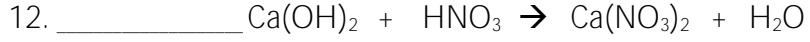
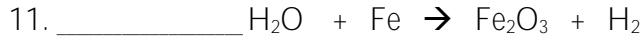
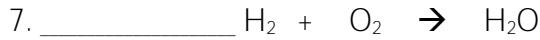
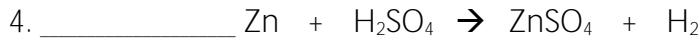
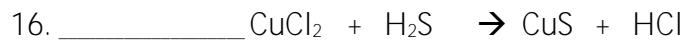


Identifying and Balancing Chemical Equations

Directions:

- Identify each of the equations below as *synthesis (S)*, *decomposition (D)*, *single replacement (SR)* or *double replacement (DR)*.
- Add coefficients to balance each of the following equations. If it is already balanced, circle the equation.





Directions: Write the chemical formulas for each of the following equations. (Don't forget to look up oxidation numbers when writing the chemical formulas.)

21. Sodium hydroxide _____

25. Lithium carbonate _____

22. Calcium fluoride _____

26. Magnesium nitrate _____

23. Aluminum oxide _____

27. Calcium phosphate _____

24. Silver nitrate _____

28. Potassium chloride _____

Directions:

- Identify each of the equations below as *synthesis (S)*, *decomposition (D)*, *single replacement (SR)* or *double replacement (DR)*.

29. _____ Silver + chlorine \rightarrow silver chloride

30. _____ Magnesium hydroxide + hydrochloric acid \rightarrow magnesium chloride + water

31. _____ Hydrochloric acid + sodium hydroxide \rightarrow sodium chloride + water

32. _____ Hydrogen + oxygen \rightarrow water

33. _____ Sodium + water \rightarrow sodium hydroxide + hydrogen

34. _____ Iron(III) oxide + carbon monoxide \rightarrow iron + carbon dioxide

35. _____ Sodium hydroxide + iron(II) chloride \rightarrow sodium chloride + iron(II) hydroxide